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It has one unpaired electron in the 3s orbital. The 3s valence atomic orbital of sodium overlaps with another such orbital of the same energy to form molecular orbitals. The atomic orbitals continue to combine in some fashion forming a band.

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$\text{Fe (s)} + \text{CuSO}_4 \text{ (aq)} \rightarrow \text{FeSO}_4 \text{ (aq)} + \text{Cu (s)}$   
 $\text{Fe (s)} + \text{FeSO}_4 \text{ (aq)} \rightarrow$  no  
reaction.  $\text{Fe (s)} + \text{Al}_2(\text{SO}_4)_3 \text{ (aq)} \rightarrow$   
no reaction.  $\text{Fe (s)} + \text{ZnSO}_4 \text{ (aq)} \rightarrow$  no  
reaction. No reaction occurs when any  
of the metals from zinc, iron,  
aluminium, and copper are placed in  
aqueous solution of  $\text{Al}_2(\text{SO}_4)_3$ .

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